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 $V = \frac{NRT}{P} = \frac{4.5 \text{ mol} \cdot 8.314 \text{ J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 300 \text{ K}}{10^5 \text{ bar}} = 1.17 \text{ m}^3$
106 volume taken up by liquid
Equating P and P_{vap} we have $x = \frac{P_{\text{vap}} V_{\text{m}}}{RT} = \frac{10^5 \text{ bar} \cdot 1.17 \text{ m}^3}{8.314 \text{ J} \cdot \text{mol}^{-1} \cdot \text{K}^{-1} \cdot 300 \text{ K}} = 4.5$

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